

Chemistry I: Midterm Review 2013

General Stuff

1. Density and formula
2. Define Accuracy and precision
3. Metric system

Matter and its changes

3. Describe the basic properties of matter in terms of mass and volume.
4. State the Law of Conservation of Matter and Energy
5. Define elements, compounds, and mixtures (provide 3 examples of each)

	definition	3 examples
element		
Mixture		
compound		

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6. List the physical and chemical properties metals, nonmetals and metalloids.

	Physical properties	Chemical properties
metals		
nonmetals		
metalloids		

7. Where are metals, nonmetals, and metalloids on periodic table?

8. Provide 5 examples of chemical change

9. 4 things to start a chemical reaction

Mid Year Problems: Show all WORK

10. Define physical and chemical changes and list 3 examples of each

11. Define a Precipitate

Chapter 3: Atomic structure

12. Define the following terms:

a) An atom is

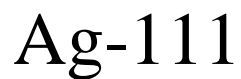
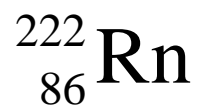
b) An Isotope is

c) The 3 basic atomic particles are

d) Atomic number is the number of

e) Atomic mass is the sum of

f) Explain the following isotopic symbol and abbreviated version.



13. Explain the significance of Ernest Rutherford's gold foil experiment

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14. Define the quantum model of the atom.

Arrangement of Electron in Atoms

15. Define the following terms:

- a) Bright line spectrum
- b) One quantum of energy
- c) One photon of light energy
- d) Valence electrons

16. Electron configuration

17. Orbital notation

The periodic law

18. the Periodic Law states that

19. Define the following terms:

- a) Ionization energy
- b) Atomic radius
- c) Electronegativity

Mid Year Problems: Show all WORK

20. Describe the trends of the following periodic properties as you go down a family / group on the Periodic Table

	Trend top to bottom	Trend left to right
Ionization energy		
Atomic radius		
Electronegativity		
Reactivity		

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Chemical bonds

21. Define the following terms:

- a) Coordinate covalent bond
- b) Double/ triple covalent bonding
- c) Polyatomic ions (radicals)
- d) Ionic bonding
- e) Define cation and anion

22. Be able to determine polarity of covalent molecules (molecular polarity)

23. Determine polar and non-polar bond using electronegativity values

Chemical Composition:

24. Define the following terms:

- a) Law of definite composition
- b) Law of multiple proportions
- c) Molar mass

a) Families b) Periods c) Alkali metals d) Alkaline earth

- 1 2

58	★	59	60	61 H	62	63	64	65	66	67	68	69	70	71
90	★	91	92	93	94	95	96	97	98	99	100	101	102	103

element	e-configuration	Lewis Dot	Ox. #
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element	e-configuration	Lewis Dot	Ox. #
A			
B			
C			
D			
E			
F			
G			
H			
I			
J			

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3. Draw the following molecules and determine their shape

Molecule	Lewis structure	Geometry
$\text{CH}_3\text{CH}_2\text{OH}$		
SCl_2		
$\text{CH}_3\text{CH}_2\text{COOH}$		

Draw examples of the following ionic bonds

Ionic Compound	Lewis drawing
LiBr	
BaCl_2	
Mg_3N_2	

Mid Year Problems: Show all WORK

Write correct formulas for the following compounds:

Potassium oxide	Carbon terta iodide	Diphosphorus pentoxide
Aluminum sulfide	Dinitrogen oxide	Silver sulfide
Copper (II) nitrate	Ammonium phosphate	Ferrous sulfate
Lead (IV) oxide	Cupric hydroxide	Stannous bicarbonate

2. Correctly name the following compounds:

$\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$	Cr_2O_3	$\text{Pb}(\text{NO}_3)_2$
Ag_2CrO_4	P_4O_{10}	NBr_3
$\text{Al}_2(\text{CrO}_4)_3$	$\text{Sn}(\text{Cr}_2\text{O}_7)_2$	N_2O_4

3. For the following calcium compounds: write the formulas and find the % of calcium in each

Calcium phosphate

Calcium nitrate

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Calculate the empirical formulas for the following

52.2% C 13.00% H 34.80% O

26.56 % K 35.41 % Cr 38.03 % O

4. Calculate the molecular formula for a compound that is 49.3% C 6.9% H and 43.8% O with a molar mass of 146 g/mol

26. Complete, identify and balance the following chemical reactions:

Composition (synthesis), Decomposition, Single Replacements, Double Replacement, Burning

